

## L-Serine methyl ester hydrochloride

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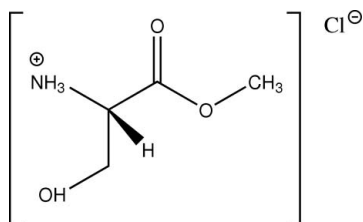
Received 30 October 2009; accepted 4 November 2009

 Key indicators: single-crystal X-ray study;  $T = 150$  K; mean  $\sigma(\text{C}-\text{C}) = 0.001$  Å;  $R$  factor = 0.022;  $wR$  factor = 0.058; data-to-parameter ratio = 34.5.

In the enantiopure crystal of the title compound,  $\text{C}_4\text{H}_{10}\text{NO}_3^+\cdot\text{Cl}^-$ , intermolecular  $\text{O}-\text{H}\cdots\text{Cl}$  and  $\text{N}-\text{H}\cdots\text{Cl}$  hydrogen bonds link the molecules into layers parallel to (001).

### Related literature

Esterification of the carboxyl group of amino acids plays an important role in the synthesis of peptides, especially due to the increased solubility in non-aqueous organic solvents, see: Bodanszky (1993). For related structures, see: Bryndal *et al.* (2006); Görbitz (1989). For the determination of the absolute structure, see: Flack & Bernardinelli (2000); Flack & Shmueli (2007); Hooft *et al.* (2008).



### Experimental

#### Crystal data

$\text{C}_4\text{H}_{10}\text{NO}_3^+\cdot\text{Cl}^-$   
 $M_r = 155.58$   
 Monoclinic,  $P2_1$   
 $a = 5.22645$  (9) Å  
 $b = 6.39388$  (14) Å  
 $c = 11.6420$  (4) Å  
 $\beta = 90.090$  (1)°

$V = 389.04$  (2) Å<sup>3</sup>  
 $Z = 2$   
 Mo  $K\alpha$  radiation  
 $\mu = 0.44$  mm<sup>-1</sup>  
 $T = 150$  K  
 $0.38 \times 0.33 \times 0.15$  mm

#### Data collection

Nonius KappaCCD diffractometer  
 Absorption correction: multi-scan  
 (*SADABS*; Sheldrick, 2002)  
 $T_{\min} = 0.78$ ,  $T_{\max} = 0.93$   
 15456 measured reflections  
 3449 independent reflections  
 3242 reflections with  $I > 2\sigma(I)$   
 $R_{\text{int}} = 0.030$

#### Refinement

$R[F^2 > 2\sigma(F^2)] = 0.022$   
 $wR(F^2) = 0.058$   
 $S = 1.07$   
 3449 reflections  
 100 parameters  
 1 restraint

H atoms treated by a mixture of independent and constrained refinement  
 $\Delta\rho_{\text{max}} = 0.25$  e Å<sup>-3</sup>  
 $\Delta\rho_{\text{min}} = -0.17$  e Å<sup>-3</sup>  
 Absolute structure: Flack (1983), 1599 Friedel pairs  
 Flack parameter: 0.00 (3)

**Table 1**

Hydrogen-bond geometry (Å, °).

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
$\text{N1}-\text{H1A}\cdots\text{Cl1}$	0.867 (16)	2.390 (16)	3.2237 (8)	161.2 (14)
$\text{N1}-\text{H1B}\cdots\text{Cl1}^{\text{i}}$	0.825 (18)	2.336 (18)	3.1563 (8)	172.7 (16)
$\text{N1}-\text{H1C}\cdots\text{Cl1}^{\text{ii}}$	0.865 (14)	2.264 (14)	3.0979 (7)	161.9 (12)
$\text{O3}-\text{H3}\cdots\text{Cl1}^{\text{iii}}$	0.847 (18)	2.261 (18)	3.1041 (7)	173.9 (15)

 Symmetry codes: (i)  $-x + 1, y - \frac{1}{2}, -z + 1$ ; (ii)  $x - 1, y, z$ ; (iii)  $x, y - 1, z$ .

Data collection: *COLLECT* (Nonius, 1999); cell refinement: *PEAKREF* (Schreurs, 2005); data reduction:  *EVAL15* (Xian *et al.*, 2006) and *SADABS* (Sheldrick, 2002); program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *PLATON* (Spek, 2009); software used to prepare material for publication: *PLATON*.

Financial assistance for this project was provided by the Dutch Organization for Scientific Research, Chemical Sciences (NWO-CW).

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: VM2012).

### References

- Bodanszky, M. (1993). *Principles of Peptide Synthesis*, 2nd ed., pp. 74–86. Berlin: Springer-Verlag.
- Bryndal, I., Jaremko, M., Jaremko, L. & Lis, T. (2006). *Acta Cryst.* **C62**, o111–o114.
- Flack, H. D. (1983). *Acta Cryst.* **A39**, 876–881.
- Flack, H. D. & Bernardinelli, G. (2000). *J. Appl. Cryst.* **33**, 1143–1148.
- Flack, H. D. & Shmueli, U. (2007). *Acta Cryst.* **A63**, 257–265.
- Görbitz, C. H. (1989). *Acta Chem. Scand.* **43**, 871–875.
- Hooft, R. W. W., Straver, L. H. & Spek, A. L. (2008). *J. Appl. Cryst.* **41**, 96–103.
- Nonius (1999). *COLLECT*. Nonius BV, Delft, The Netherlands.
- Schreurs, A. M. M. (2005). *PEAKREF*. Utrecht University, The Netherlands.
- Sheldrick, G. M. (2002). *SADABS*. University of Göttingen, Germany.
- Sheldrick, G. M. (2008). *Acta Cryst.* **A64**, 112–122.
- Spek, A. L. (2009). *Acta Cryst.* **D65**, 148–155.
- Xian, X., Schreurs, A. M. M. & Kroon-Batenburg, L. M. J. (2006). *Acta Cryst.* **A62**, s92.

**supplementary materials**

*Acta Cryst.* (2009). E65, o3026 [ doi:10.1107/S1600536809046480 ]

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A. Schouten and M. Lutz

### Comment

Esterification of the carboxyl group of amino acids plays an important role in the synthesis of peptides, especially due to the increased solubility in non-aqueous organic solvents (Bodanszky, 1993). The synthesis of methyl esters is straightforward and can be performed by the reaction of HCl gas with a suspension of the amino acid in methanol. In this reaction the hydrochloride of the amino acid methyl ester is obtained, which is the subject of the present study.

A molecular plot of the title compound (I) is shown in Fig. 1. As a consequence of the protection of the carboxyl group, only the amino group is available as protonation site. The O1—C1—C2—N1 and C4—O1—C1—C2 torsion angles of  $-175.99(7)$  and  $179.72(7)^\circ$  indicate an extended structure of the backbone. Nevertheless, we do not see a stabilization of this extended structure by an intramolecular hydrogen bond between the ammonium group and O2. Such a stabilization would require a N—H bond in the plane of O2—C1—C2—N1, which is not the case here. Similar extended structures are also found in the methyl ester hydrochlorides of L-cysteine (Görbitz, 1989) and L-tyrosine (Bryndal *et al.*, 2006). Bond lengths and angles in (I) are as expected (Table 1).

With three H-atoms at N1 and one H-atom at O3 the molecule has four hydrogen bond donors. Three oxygen atoms and the chloride anion could act as hydrogen bond acceptors. In fact, only the chloride is used as a hydrogen bond acceptor, here (Table 2). This results in an infinite two-dimensional hydrogen bonding network parallel to the *ab* plane, as shown in Fig. 2. In the *c* direction the hydrogen-bonded planes are separated by hydrophobic OCH<sub>3</sub> groups. Interestingly, the two-dimensional motif is also reflected in the morphology of the crystal, where (001) has the smallest dimension.

Despite a  $\beta$ -angle of  $90.090(1)^\circ$  there is no orthorhombic symmetry in this crystal structure. The  $R_{\text{int}}$  value for Laue-symmetry *mmm* is 31% compared to 2% for *2/m*. We also did not find indications for pseudo-orthorhombic twinning. The reflections in the diffraction images were not split, and an analysis of the  $F_o/F_c$  listing with the TWINROT/MAT routine of *PLATON* (Spek, 2009) did not suggest the presence of twinning.

Because (I) is derived from enantiopure L-serine, the absolute configuration was known in advance. Nevertheless, the presence of chloride provides enough enantiomorph distinguishing power (Friedif = 123, Flack & Shmueli, 2007) to allow a reliable experimental confirmation of the absolute structure. This was done using the Flack parameter (Flack, 1983), which resulted in  $x = 0.00(3)$ , and the Hooft parameter (Hooft *et al.*, 2008), which resulted in  $y = 0.005(15)$ . As expected, the standard uncertainty of the Hooft parameter is significantly lower than in the Flack parameter, but both parameters confirm the correct absolute structure of (I).

### Experimental

0.4 g of L-serine methyl ester hydrochloride (obtained commercially from Aldrich) was dissolved in 10 ml absolute ethanol followed by slow evaporation at room temperature. Single crystals suitable for X-ray diffraction were obtained after a final seeding step by adding a tiny amount of solid starting material.

## Refinement

All H atoms were located in difference Fourier maps. H atoms bonded to N and O atoms were refined freely with isotropic displacement parameters. H(C) atoms were refined using a riding model (including free rotation of the methyl substituents), with C—H = 0.95–1.00 Å and with  $U_{\text{iso}}(\text{H}) = 1.2$  (1.5 for methyl groups) times  $U_{\text{eq}}(\text{C})$ .

Friedel pairs were kept separate during the refinement and the Flack parameter was included in the least-squares matrix using TWIN/BASF instructions in *SHELXL97*. This has been shown to reduce the uncertainty of the Flack parameter compared to the hole-in-one algorithm (Flack & Bernardinelli, 2000).

## Figures

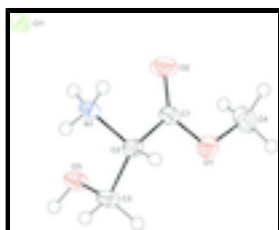


Fig. 1. Molecular structure of (I), with atom labels and 50% probability displacement ellipsoids for non-H atoms.

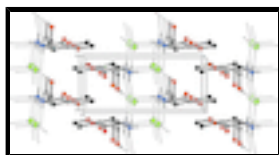


Fig. 2. The packing of (I), viewed down the *b* axis, showing sheets running parallel to (001) with molecules connected by O—H...Cl and N—H...Cl hydrogen bonds (dashed lines). H atoms not involved in hydrogen bonding have been omitted for clarity.

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### Crystal data

$\text{C}_4\text{H}_{10}\text{NO}_3^+\cdot\text{Cl}^-$	$F_{000} = 164$
$M_r = 155.58$	$D_x = 1.328 \text{ Mg m}^{-3}$
Monoclinic, $P2_1$	Mo $K\alpha$ radiation, $\lambda = 0.71073 \text{ \AA}$
Hall symbol: P 2yb	Cell parameters from 14244 reflections
$a = 5.22645 (9) \text{ \AA}$	$\theta = 1.8\text{--}35.0^\circ$
$b = 6.39388 (14) \text{ \AA}$	$\mu = 0.44 \text{ mm}^{-1}$
$c = 11.6420 (4) \text{ \AA}$	$T = 150 \text{ K}$
$\beta = 90.090 (1)^\circ$	Irregular plate, colourless
$V = 389.044 (16) \text{ \AA}^3$	$0.38 \times 0.33 \times 0.15 \text{ mm}$
$Z = 2$	

### Data collection

Nonius KappaCCD diffractometer	3449 independent reflections
Radiation source: rotating anode	3242 reflections with $I > 2\sigma(I)$
Monochromator: graphite	$R_{\text{int}} = 0.030$

$T = 150$  K  $\theta_{\max} = 35.0^\circ$   
 $\varphi$  and  $\omega$  scans  $\theta_{\min} = 1.8^\circ$   
 Absorption correction: multi-scan (SADABS; Sheldrick, 2002)  $h = -8 \rightarrow 8$   
 $T_{\min} = 0.78$ ,  $T_{\max} = 0.93$   $k = -10 \rightarrow 10$   
 15456 measured reflections  $l = -18 \rightarrow 18$

### Refinement

Refinement on  $F^2$  Hydrogen site location: difference Fourier map  
 Least-squares matrix: full H atoms treated by a mixture of independent and constrained refinement  
 $R[F^2 > 2\sigma(F^2)] = 0.022$   $w = 1/[\sigma^2(F_o^2) + (0.0329P)^2 + 0.0147P]$   
 where  $P = (F_o^2 + 2F_c^2)/3$   
 $wR(F^2) = 0.058$   $(\Delta/\sigma)_{\max} = 0.002$   
 $S = 1.07$   $\Delta\rho_{\max} = 0.25 \text{ e } \text{\AA}^{-3}$   
 3449 reflections  $\Delta\rho_{\min} = -0.17 \text{ e } \text{\AA}^{-3}$   
 100 parameters Extinction correction: none  
 1 restraint Absolute structure: Flack (1983), 1600 Friedel pairs  
 Primary atom site location: structure-invariant direct methods Flack parameter: 0.00 (3)  
 Secondary atom site location: difference Fourier map

### Special details

**Geometry.** All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

**Refinement.** Refinement of  $F^2$  against ALL reflections. The weighted  $R$ -factor  $wR$  and goodness of fit  $S$  are based on  $F^2$ , conventional  $R$ -factors  $R$  are based on  $F$ , with  $F$  set to zero for negative  $F^2$ . The threshold expression of  $F^2 > \sigma(F^2)$  is used only for calculating  $R$ -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement.  $R$ -factors based on  $F^2$  are statistically about twice as large as those based on  $F$ , and  $R$ -factors based on ALL data will be even larger.

### Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{\AA}^2$ )

	$x$	$y$	$z$	$U_{\text{iso}}^*/U_{\text{eq}}$
O1	0.19488 (14)	0.39944 (12)	0.93010 (6)	0.03194 (14)
O2	0.32242 (18)	0.69384 (13)	0.84254 (7)	0.04120 (18)
O3	0.57878 (12)	0.22667 (10)	0.72179 (7)	0.02923 (13)
H3	0.644 (3)	0.119 (3)	0.6913 (14)	0.045 (4)*
N1	0.25944 (14)	0.54258 (11)	0.63222 (7)	0.02279 (12)
H1A	0.411 (3)	0.597 (3)	0.6391 (12)	0.037 (4)*
H1B	0.254 (3)	0.473 (3)	0.5725 (15)	0.045 (4)*
H1C	0.149 (2)	0.643 (2)	0.6261 (11)	0.029 (3)*
C1	0.24364 (16)	0.51767 (14)	0.83958 (7)	0.02379 (14)
C2	0.18095 (14)	0.40362 (12)	0.72813 (7)	0.02175 (13)

## supplementary materials

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H2	-0.0084	0.3827	0.7239	0.026*
C3	0.31005 (16)	0.19224 (13)	0.71800 (8)	0.02470 (14)
H3A	0.2571	0.1003	0.7821	0.030*
H3B	0.2616	0.1242	0.6447	0.030*
C4	0.2501 (2)	0.49492 (19)	1.04131 (9)	0.0378 (2)
H4A	0.4322	0.5307	1.0456	0.057*
H4B	0.2079	0.3961	1.1028	0.057*
H4C	0.1473	0.6220	1.0502	0.057*
Cl1	0.77391 (3)	0.81723 (3)	0.608559 (15)	0.02488 (4)

### Atomic displacement parameters ( $\text{\AA}^2$ )

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
O1	0.0416 (3)	0.0290 (3)	0.0252 (3)	-0.0042 (3)	0.0023 (2)	0.0007 (2)
O2	0.0668 (5)	0.0250 (3)	0.0317 (3)	-0.0129 (4)	-0.0020 (3)	-0.0044 (3)
O3	0.0229 (3)	0.0219 (3)	0.0429 (4)	0.0033 (2)	-0.0034 (2)	-0.0061 (2)
N1	0.0220 (3)	0.0201 (3)	0.0263 (3)	0.0021 (2)	-0.0026 (2)	-0.0013 (2)
C1	0.0236 (3)	0.0218 (3)	0.0260 (3)	0.0014 (3)	0.0001 (3)	-0.0023 (3)
C2	0.0187 (3)	0.0196 (3)	0.0270 (3)	-0.0016 (3)	-0.0014 (2)	-0.0012 (3)
C3	0.0250 (3)	0.0168 (3)	0.0323 (4)	-0.0016 (3)	-0.0027 (3)	-0.0025 (3)
C4	0.0485 (6)	0.0403 (5)	0.0248 (4)	0.0096 (4)	0.0003 (4)	-0.0030 (4)
Cl1	0.02321 (7)	0.02180 (7)	0.02965 (8)	0.00378 (7)	0.00166 (5)	0.00182 (8)

### Geometric parameters ( $\text{\AA}$ , $^\circ$ )

O1—C1	1.3221 (11)	C1—C2	1.5235 (12)
O1—C4	1.4598 (13)	C2—C3	1.5153 (12)
O2—C1	1.1998 (11)	C2—H2	1.0000
O3—C3	1.4223 (10)	C3—H3A	0.9900
O3—H3	0.847 (18)	C3—H3B	0.9900
N1—C2	1.4852 (11)	C4—H4A	0.9800
N1—H1A	0.867 (16)	C4—H4B	0.9800
N1—H1B	0.825 (18)	C4—H4C	0.9800
N1—H1C	0.865 (14)		
C1—O1—C4	115.45 (8)	C3—C2—H2	108.5
C3—O3—H3	104.9 (11)	C1—C2—H2	108.5
C2—N1—H1A	115.1 (10)	O3—C3—C2	107.42 (6)
C2—N1—H1B	107.6 (12)	O3—C3—H3A	110.2
H1A—N1—H1B	108.9 (14)	C2—C3—H3A	110.2
C2—N1—H1C	108.6 (9)	O3—C3—H3B	110.2
H1A—N1—H1C	108.6 (14)	C2—C3—H3B	110.2
H1B—N1—H1C	107.7 (14)	H3A—C3—H3B	108.5
O2—C1—O1	125.48 (9)	O1—C4—H4A	109.5
O2—C1—C2	123.17 (8)	O1—C4—H4B	109.5
O1—C1—C2	111.33 (7)	H4A—C4—H4B	109.5
N1—C2—C3	110.58 (7)	O1—C4—H4C	109.5
N1—C2—C1	107.14 (6)	H4A—C4—H4C	109.5
C3—C2—C1	113.45 (7)	H4B—C4—H4C	109.5

N1—C2—H2	108.5		
C4—O1—C1—O2	-1.64 (14)	O2—C1—C2—C3	127.65 (10)
C4—O1—C1—C2	179.72 (7)	O1—C1—C2—C3	-53.67 (9)
O2—C1—C2—N1	5.33 (11)	N1—C2—C3—O3	59.87 (9)
O1—C1—C2—N1	-175.99 (7)	C1—C2—C3—O3	-60.53 (9)

*Hydrogen-bond geometry (Å, °)*

<i>D</i> —H $\cdots$ <i>A</i>	<i>D</i> —H	H $\cdots$ <i>A</i>	<i>D</i> $\cdots$ <i>A</i>	<i>D</i> —H $\cdots$ <i>A</i>
N1—H1A $\cdots$ C11	0.867 (16)	2.390 (16)	3.2237 (8)	161.2 (14)
N1—H1B $\cdots$ C11 <sup>i</sup>	0.825 (18)	2.336 (18)	3.1563 (8)	172.7 (16)
N1—H1C $\cdots$ C11 <sup>ii</sup>	0.865 (14)	2.264 (14)	3.0979 (7)	161.9 (12)
O3—H3 $\cdots$ C11 <sup>iii</sup>	0.847 (18)	2.261 (18)	3.1041 (7)	173.9 (15)

Symmetry codes: (i)  $-x+1, y-1/2, -z+1$ ; (ii)  $x-1, y, z$ ; (iii)  $x, y-1, z$ .

Fig. 1

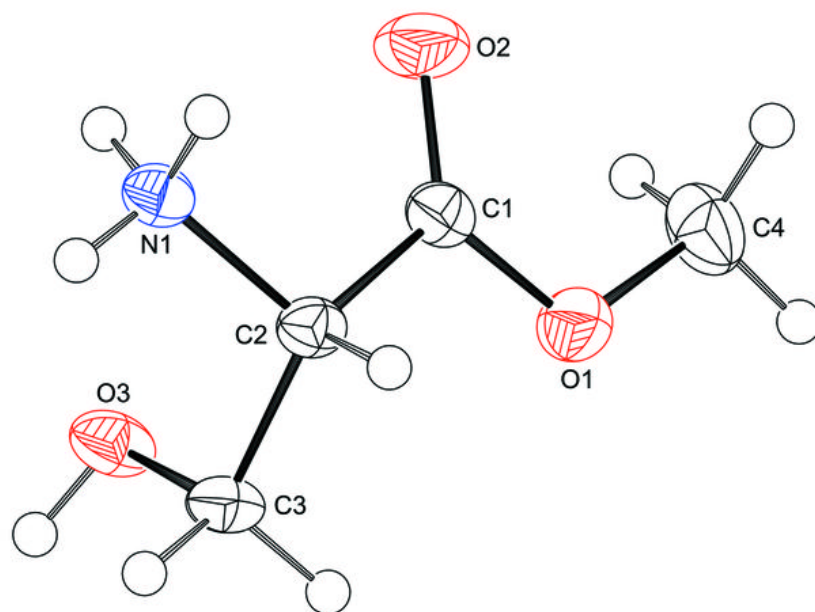




Fig. 2

